

Some basic concepts of chemistry neet questions pdf

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Students can check JEE NEET Chemistry Question Bank for Basic Concepts of Chemistry below and to check the question bank of Chemistry's other topics, one can use the button given at the bottom of this page.JEE NEET Chemistry Question Bank - Basic Concepts of Chemistry Download Chemistry Question BankBiologyMathsPhysicsTo get fastest exam alerts and government job alerts in India, join our Telegram channel. Question 1: Question 2: Question 3: 74.5 g of a metallic chloride contain 35.5 g of chlorine. The equivalent weight of the metal is (a) 19.5 (b) 35.5 (c) 39.0 (d) 78.0 Question 4: Question 5: The vapour density of a volatile chloride of a metal is 95 and the specific heat of the metal is 0.13 cal/g. The equivalent weight of the metal will be approximately (a) 6 (b) 12 (c) 18 (d) 49 Question 6: Question 7: Question 8: Question 9: What is the percentage of cation in ammonium dichromate? (a) 14.29% (b) 80% (c) 50.05% (d) 20.52% Question 10: Question 11: Question 12: Question 13: Question 14: One mole of chlorine combines with certain weight of a metal giving 111 g of its chloride. The same amount of metal can displace 2 g of hydrogen from an acid. The atomic weight of the metal is (a) 40 (b) 20 (c) 80 (d) None of these Question 15: 1.520 g of hydroxide of a metal on ignition gave 0.995 g of oxide. The equivalent weight of metal is (a) 1.52 (b) 0.995 (c) 190 (d) 9 Question 16: The mass of carbon anode consumed (given only C02) in production of 270 kg of aluminium metal from bauxite by the Hall process is (atomic mass of Al =27) (a) 180 kg (b) 270 kg (c) 540 kg (d)90kg Question 17: What volume of a solution of hydrochloric acid containing 73g of acid per litre would sufficient for exact neutralisation of sodium hydroxide obtained by allowing 0.46 g of metallic sodium to act upon water? (Cl = 35.5, Na = 23.0, O = 16) (a)10mL (b)20mL (c) 30 mL (d) 40 mL Question 18: 0.5 g of fuming H2SO4(oleum) is diluted with water. This solution is completely neutralised by 26.7 mL of 0. 4 N NaOH. The percentage of free SO3 in the sample is (a) 30.6% (b) 40.6% (c) 20.6% (d) 50% Question 19: Question 20: Question 21: Question 22: Question 23: Question 24: Question 25: An aqueous solution of glucose is 10% in strength. The volume in which 1g mole of it is dissolved will be (a) 18 L (b) 92 L (c)0.9L (d)1.8L Question 26: Question 27: 0.45 g acid of molecular weight 90 was neutralise by 20 mL of 0.5 N KOH. The basicity of acid is (a) 2 (b) 4 (c) 1 (d) 3 Question 28: Direction (Q. NO. 29) In the following question more than one of the answers may be correct. Select the correct answers and mark it according to the codes. Codes (a) I, II and III are correct (b) I and II are correct (c) II and IV are correct (d) I and III are correct Question 29: Direction (Q. NOS. 31-34) Each of these questions contains two statements - Statement I and II. Each of these questions also has four alternative choices, only one of which is the correct answer. You have to select one of the codes (a), (fe), (c) and (d) given below. (a) Statement I is true, Statement II is true; Statement II is a correct explanation for Statement I (b) Statement I is true, Statement II is true; Statement II is not a correct explanation for Statement I (c) Statement I is true, Statement II is false (d) Statement I is false, Statement II is true Question 30: Question 31: Question 32: Statement I The molality of the solution does not change with change in temperature. Statement II The molality of the solution is expressed in units of moles per 1000 g of solvent. Question 33: Normality formula in chemistry is one of the important term used to measure the concentration of a solution. Question 34: Statement I Normality and molarity can be calculated from each other. Statement II Normality is equal to the product of molarity and n. Question 35: Question 36: When 22.4 L of H2(g) is mixed with 11.2 L of Cl2(g) each at STP, the moles of HCl (g) formed is equal to (a)1 mole of HCl (g) (b)2 moles of HCl (g) (c) 0.5 mole of HCl (g) (d)5 mole of HCl (g) Question 37: Question 38: Question 39: The decomposition of a certain mass of CaCO3 gave 11.2 dm3 of CO 2 gas at STP. The mass of KOH required to completely neutralise the is [AIIMS 2012] (a)56 g (b) 28 g (c) 42 g (d) 20 g Question 40: A metal oxide has the formula A2O3. It can be reduced by hydrogen to give free metal and water. 0.1596 g of this metal oxide requires 6 mg of hydrogen for complete reduction. What is the atomic weight of metal? (AI3) (b) 57.3 (c) 55.8 (d) 59.3 Question 41: The weight of iron which will be converted into its oxide (Fe3O4) by the action of 18 g of steam on it will be (atomic weight of Fe = 56) (a) 168 g (b)84g (c) 42 g (d) 21 g Question 42: Statement I Equal moles of different substances contain same number of constituent particles. Statement II Equal weights of different substances contain the same number of constituent particles. (a) Statement I is true, Statement II is true; Statement II is a correct explanation for Statement I (b) Statement I is true, Statement II is true; Statement II is not a correct explanation for Statement I (c) Statement I is true, Statement II is false (d) Statement I is false, Statement II is true Question 43: Question 44: Question 44: Question 45: Answers: Hints and Solutions ChemistryPhysicsBiology Guys, does anyone know the answer? get some basic concepts of chemistry neet questions from screen.Some Basic Concepts Of Chemistry Chemistry NEET Practice Questions, MCQs, Past Year Questions (PYQs), NCERT Questions, Question Bank, Class 11 and Class 12 Questions, and PDF solved with answersSome Basic Concepts Of Chemistry Chemistry Practice questions, MCQs, Past Year Questions (PYQs), NCERT Questions, Question Bank, Class 11 and Class 12 Questions, NCERT Exemplar Questions and PDF Questions with answers, solutions, explanations, NCERT reference and difficulty levelx 8527521718 Online Support Signup | LoginSelect Question Set: Recommended MCQs - 118 Questions Recommended PYQs (STRICTLY NCERT Based) NCERT Solved Examples Based MCQs NCERT Exercise Based MCQs NCERT Exemplar (Objective) Based MCQs AR & Other Type MCQs Past Year (2019 onward - NTA Papers) MCQs Past Year (2016 - 2018) MCQs Past Year (2006 - 2015) MCQs1 (current) 2 3 4 Q. No.A mixture of gases contains H2 and O2 gases in the ratio of 1:4 (w/w). What is the molarratio of the two gases in the mixture?1. 1:4 2. 3:4 3. 16:1 4. 2:1 Q1:59% From NCERT NEET - 2015Subtopic: Moles, Atoms & Electrons |If the Avogadro number NA, is changed from 6.022 x 1023 mol-1 to 6.022 x 1020 mol-1 thiswould change1. The definition of mass in units of grams2. The mass of one mole of carbon3. The ratio of chemical species to each other in a balanced equation4. The ratio of elements to each other in a compoundQ2:70% From NCERT NEET - 2015Subtopic: Moles, Atoms & Electrons |20.0 g of a magnesium carbonate sample decomposes on heating to give carbon dioxideand 8.0 g magnesium oxide. What will be the percentage purity of magnesium carbonateln the sample? (Atomic weight of Mg = 24)1. 75 2. 96 3. 60 4. 84 Q3: 52% NEET - 2015Subtopic: Equation Based Problem |When 50 mL of a 16.9 % (w/v) solution of AgNO3 is mixed with 50 mL of 5.8% (w/v) NaCl solution, then the mass of precipitate formed is(Ag = 107.8, N = 14, O = 16, Na= 23,Cl=35.5)1. 28 g 2. 3.5 g 3. 7 g 4. 14 g Q4:50% From NCERT NEET - 2015Subtopic: Equation Based Problem |The number of water molecules is maximum in -1. 18 molecules of water2. 1.8 g of water3. 18 g of water4. 18 moles of waterQ5:72% From NCERT NEET - 2015Subtopic: Moles, Atoms & Electrons |25.3 g of Sodium carbonate Na2CO3 is dissolved in enough water to make 250 mL ofsolution. If sodium carbonate dissociates completely, molar concentration of sodium ion,Na+ and carbonate ionCO 2 – 3 CO32-are respectively (Molar mass of Na2CO3 = 106 g mol-1)1. 0.955 M and 1.910 M2. 1.910 M and 0.955 M3. 1.90 M and 1.910 M4. 0.477 M and 0.477 MQ6:65% From NCERT NEET - 2010Subtopic: Concentration Based Problem |The number of atoms in 0.1 mole of a triatomic gas is(NA=6.02 x1023mol-1)1. 6.026 x 1022 2. 1.806x1023 3. 3.600 x 1023 4. 1.800 x 1022 Q7:69% From NCERT NEET - 2010Subtopic: Moles, Atoms & Electrons |10 g of hydrogen and 64 g of oxygen were filled in a steel vessel and exploded. The amount of water produced in this reaction will be:1. 2 mol 2. 3 mol 3. 4 mol 4. 1 mol Q8:65% From NCERT NEET - 2009Subtopic: Equation Based Problem |How many moles of lead (II) chloride will be formed from a reaction between 6.5g of PbO and 3.2g of HCl?(1) 0.044 (2) 0.333 (3) 0.011 (4) 0.029 Q9:From NCERT NEET - 2008Subtopic: Equation Based Problem |Concentrated aqueous sulphuric acid is 98%H 2 SO 4 by mass and has a density of 1.80 g mL-1. Volume of acid required to make one litre of 0.1 M H2SO4 solution is :(1) 11.10 mL (2) 16.65 mL (3) 22.20 mL (4) 5.55 mL Q10:56% From NCERT NEET - 2007Subtopic: Concentration Based Problem |Select Question Set: Recommended MCQs - 118 Questions Recommended PYQs (STRICTLY NCERT Based) NCERT Solved Examples Based MCQs NCERT Exercise Based MCQs NCERT Exemplar (Objective) Based MCQs AR & Other Type MCQs Past Year (2019 onward - NTA Papers) MCQs Past Year (2016 - 2018) MCQs Past Year (2006 - 2015) MCQs1 (current) 2 3 4 Q. No. x[ ] [ ] [ ] : www.neetprep.comMCQs on Some Basic Concept of Chemistry for NEET 2022Get access to MCQs on Some Basic Concept of Chemistry along with answers at BYJU'S. These questions are aligned with the NEET syllabus and help you better prepare for NEET 2022.NEETNEET MCQNEET Chemistry McqMCQs on Some Basic Concept Of ChemistryPrevious NextMCQs on Some Basic Concepts of Chemistry for NEETMCQs on Some Basic Concepts of Chemistry for NEET NEET Questions on Some Basic Concepts of ChemistryChemistry is the study of composition, properties and interaction of matter. Chemistry interlinks other branches of science. Principles of chemistry are applied widely. There are some basic concepts which help us to understand chemistry better. These are state of a matter, nature, physical and chemical properties, moles and molar mass, composition and various basic laws like Avogadro's law, Dalton's theory, etc. Below listed are some questions related to basic concepts of chemistry important for NEET.1. Elements X and Y combine to form two compounds XY and X2Y. Find the atomic weight of X and Y, if the weight of 0.1 moles of XY is 10g and 0.05 moles of X2Y is 9g(a) 30, 20 (b) 80, 20 (c) 60, 40 (d) 20, 30Answer: (b)2. Which one will have maximum numbers of water molecules?(a) 18 molecules of water(b) 1.8 grams of water(c) 18 grams of water(d) 18 moles of waterAnswer: (d)3. The number of atoms present in 0.1 moles of a triatomic gas is(a) 1.806 × 1023 (b) 1.806 × 1022 (c) 3.600 × 1023 (d) 6.026 × 1022Answer: (a)4. Find the volume of O2 required to burn 1 L of propane completely, measured at 0°C temperature and 1 atm pressure(a) 10 L (b) 7 L (c) 6 L (d) 5 LAnswer: (d)5. A gas X has Cp and Cv ratio as 1.4, at NTP 11.2 L of gas X will contain \_\_\_\_\_ number of atoms(a) 1.2 × 1023 (b) 3.01 × 1023 (c) 2.01 × 1023 (d) 6.02 × 1023Answer: (d)6. Which of the species is not paramagnetic?(a) As+ (b) Cl- (c) Ne2+ (d) Be+Answer: (b)7. Pressure has the same dimension as \_\_\_\_\_ (a) energy per unit volume(b) energy/(c) force per unit volume(d) forceAnswer: (a)8. A container has an equal mass of H2, O2 and CH4 at 27°C, the ratio of their volume is(a) 16:8:1 (b) 8:1:2 (c) 16:1:2 (d) 8:16:1Answer: (c)9. There are two chlorides of sulphur S2Cl2 and SCl2. What is the equivalent mass of S in SCl2(a) 64.8 g/mole (b) 32 g/mole (c) 16 g/mole (d) 8 g/moleAnswer: (c)10. Boron exists as two stable isotopes, 10B(19%) and 11B(81%). Find out the average atomic weight of boron in the periodic table(a) 10.0 (b) 11.2 (c) 10.2 (d) 10.8Answer: (d)11. The mass of carbon-12 atom considered in the definition of a mole is(a) 0.12g (b) 0.012 kg (c) 120 mg (d) None of theseAnswer: (b)12. Which of the following is a pair of physical and chemical property respectively of a substance(a) density and acidity(b) basicity and colour(c) colour and density(d) acidity and combustibilityAnswer: (a)13. The normality of 0.3M phosphorous acid is(a) 0.9 (b) 0.1 (c) 0.6 (d) 0.3Answer: (c)14. 3g of an oxide of a metal is converted into chloride and it yielded 5g of chloride. Find the equivalent weight of the metal (a) 20 (b) 12 (c) 3.325 (d) 33.25Answer: (d)15. 0.32 gm of a metal on treatment with an acid gave 112mL of hydrogen at STP. Calculate the equivalent weight of the metal. (a) 24 (b) 11.2 (c) 32 (d) 58Answer: (c)16. Which of the following is not a homogeneous mixture?(a) Brass (b) Air (c) Smoke(d) Aqueous solution of sugarAnswer: (c)8,83,456Don't miss:NEET Chemistry MCQsChemistry syllabus for NEETThe reaction of tert butyl bromide with sodium methoxide produces mainly – Find the answerRecommended NEET Chemistry MCQsMCQs on Organic Chemistry MCQs on P Block ElementsMCQs on States Of Matter MCQs on Redox ReactionMCQs on Chemical Equilibrium MCQs on ThermodynamicsMCQs on Periodic Classification of Elements MCQs on s Block ElementsMCQs on Chemical Bonding Molecular Structure MCQs on Structure of AtomMCQs on Environmental Chemistry MCQs on HydrocarbonsNEET Related LinksNEET Rank Predictor NEET 2018 Question PaperTop Medical Colleges In India NEET SyllabusNEET Question Paper NEET RegistrationNEET Hall Ticket NEET Results DateNEET Sample Paper NEET Books[ ] [ ] [ ] : byjus.comNEET Some Basic Concepts of Chemistry Important QuestionsSome Basic Concepts of Chemistry NEET Important Questions free pdf download. 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The pdf constitutes important questions along with step-wise solutions for each question to help students understand the concepts easily. 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T&C Apply\*Latest Vedantu courses for youGrade 11 Science | ALLBOARDS | JEE | EnglishJEE 2-Year (2022-24)Academic year 2022-24ENGLISHUnlimited access till final school examJEE Full course1000+ chapter specific short coursesPhysics Chemistry Maths ₹16,500 for 2 years Select and buyNEET 2022 Exam HighlightsThe following table gives you an overview of the NEET 2022 Exams.Particulars Details Exam NameNational Eligibility cum Entrance Test (NEET)Conducting BodyNational Testing Agency (NTA)Category Type UndergraduateFrequency of NEET ExamOnce in a Year Duration of Exam3 hours and 20 minutesTotal Number of Questions in NEET200 Questions NEET Total Marks 720 Marks Total Sections2 sections; Section A and Section BMarking Scheme+ 4 for every correct answer, -1 for an incorrect answer, 0 marks for unattempted questionsType of QuestionsMultiple Choice Questions (MCQs)Exam ModeOffline (Pen and Paper Test)NEET Language options13 languages: English, Hindi, Assamese, Bengali, Gujarati, Kannada, Malayalam, Marathi, Odia, Punjabi, Tamil, Telugu, and Urdu.NEET 2022 Exam PatternThe following table shows the details for the number of questions that will be asked in the NEET 2022 exam in each subject including Physics, Chemistry, and Biology (Botany & Zoology).Section No. of QuestionNEET Marks DistributionChemistrySection A: 35 QuestionsSection B: 15 Questions (Only 10 have to be attempted)Section A: 140 Section B: 40 Total: 180 PhysicsSection A: 35 QuestionsSection B: 15 Questions (Only 10 have to be attempted)Section A: 140 Section B: 40 Total: 180 BotanySection A: 35 QuestionsSection B: 15 Questions (Only 10 have to be attempted)Section A: 140 Section B: 40 Total: 180 ZoologySection A: 35 QuestionsSection B: 15 Questions (Only 10 have to be attempted)Section A: 140 Section B: 40 Total: 180 TotalTotal Number of Questions: 180Total Marks: 720NEET 2022 Chemistry: Chapter- Wise WeightageThe following table shows the different chapters students need to prepare for the NEET 2022 Chemistry exam along with the weightage of each chapter.Chapter Name No. of Questions WeightageSome Basic Concepts Of Chemistry1 2.0%Structure of an Atom1 3.0%Periodic Table and Periodicity in Properties3 6.0%Chemical Bonding and Molecular Structure4 9.0% Solid State 1 3.0.% States of Matter 1 2.0% Thermodynamics 1 3.0% Solutions 2 4.0%Organic Chemistry - Some Basic Principles And Techniques2 5.0% Electrochemistry 1 3.0% Chemical Kinetics 2 4.0%[ ] [ ] [ ] : www.vedantu.com







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